LO: Students will be able to complete all steps of a stoichiometry problem.

DOL: Students will be able to answer stoichiometry questions correctly at least 4/5 times.

Steps to a complete stoichiometric problem from reactants:

1) Determine the type of reaction and products

2) Write and balance a chemical equation

3) Determine the limiting reagent

4) Calculate mols of products formed and convert to grams

5) Calculate amount of non-limiting reagent that is remaining

6) Check work using the Law of Conservation of Mass

DO NOT WRITE THIS SLIDE

*these problems are long and take up a lot of space. We will only have time to do 2 in class. You will want to write down each question on a separate sheet of paper so that you have enough space to solve them.

Given 132.4 g of lithium hydroxide and 203.2 g of sulfuric acid, determine the limiting reagent, the mass of the products, and the mass of the remaining reactant. What is the percent yield if 215 grams of lithium sulfate is produced?

Given 3.45 g of carbon tetraflouride and 12.1 g of bromine, determine how much of each product will be formed and how much of the reactants will remain.