LO: Students will be able to balance chemical equations.

DOL: Students will be able to correctly balance at least 4/5 chemical equations.

Find the molar mass of each of the following compounds.

$2H_2 + O_2 \longrightarrow 2H_2O$

reactants products

Multiply each molar mass by the number in front of it. What do you notice about the total masses on each side of the arrow? Compare the masses of the reactants to the products...

Dec 7-8:34 AM

Chemical equations must be balanced before you can do anything with them.

Coefficients represent the molar ratios of the compounds in an equation.

You must have the same number of mols of EACH element in order to be balanced.

Dec 7-8:39 AM

Using a table to balance an equation:

$$H_2SO_4 + NaOH \longrightarrow H_2O + Na_2SO_4$$

Leave about 6 lines blank in your notebook to work this out in class

Dec 7-8:41 AM

Potassium reacts with water yielding potassium hydroxide and hydrogen.

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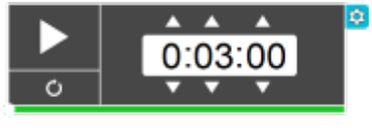
Dec 4-12:33 PM

Balance the equation.

$$Al + N_2 \rightarrow AlN$$

Leave about 6 lines blank in your notebook to work this out in class



Dec 7-8:43 AM

Balance the equation.

$$Fe + O_2 \rightarrow Fe_3O_4$$

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Dec 7-8:43 AM

Balance the equation.

$$CaCO_3 \rightarrow CaO + CO_2$$

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Balance the equation.

$$NH_4NO_3 \rightarrow N_2O + H_2O$$

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Balance the equation.

$$KI + Cl_2 \rightarrow KCl + I_2$$

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Balance the equation.

$$Pb(NO_3)_2 + HCl \rightarrow PbCl_2 + HNO_3$$

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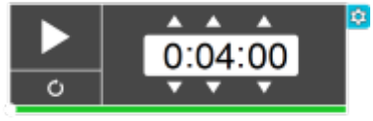
Balance the equation.



$$BaO_2 \rightarrow BaO + O_2$$

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Balance the equation.

$$Al + H_2SO_4 \rightarrow Al_2(SO_4)_3 + H_2$$

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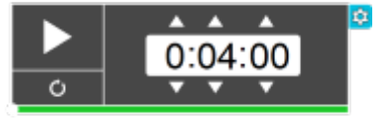
Balance the equation.



$$CH_4 + Cl_2 \rightarrow CHCl_3 + HCl$$

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Balance the equation.

$$MgCl_2 + NaOH \rightarrow Mg(OH)_2 + NaCl$$

Leave about 6 lines blank in your notebook to work this out in class

Dec 7-8:43 AM