

Naming Covalent Compounds

Since the ionic charges of the individual atoms do not matter, you must state how many of each atom you have.

However, you never start a molecule's name with mono

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Prefixes for naming molecules

1 = mono

6 = hexa

2 = di

7 = hepta

3 = tri

8 = octa

4 = tetra

9 = nona

5 = penta

10 = deca

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Name the following molecular compounds



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Naming Acids and Bases

Acids almost always begin with a H

Bases almost always end in an OH

Acids get their name based on the anion that the hydrogen is bonded to.

Bases get their names based on the cation

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Some common acids and their names

anion	anion name	acid	acid name
Cl^-	chloride ion	HCl	hydrochloric acid
CO_3^{2-}	carbonate ion	H_2CO_3	carbonic acid
NO_2^-	nitrite ion	HNO_2	nitrous acid
NO_3^-	nitrate ion	HNO_3	nitric acid
SO_3^{2-}	sulfite ion	H_2SO_3	sulfurous acid
SO_4^{2-}	sulfate ion	H_2SO_4	sulfuric acid
CH_3COO^-	acetate ion	CH_3COOH	acetic acid

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Bases are Ionic Compounds, so name them just like you would an ionic compound.

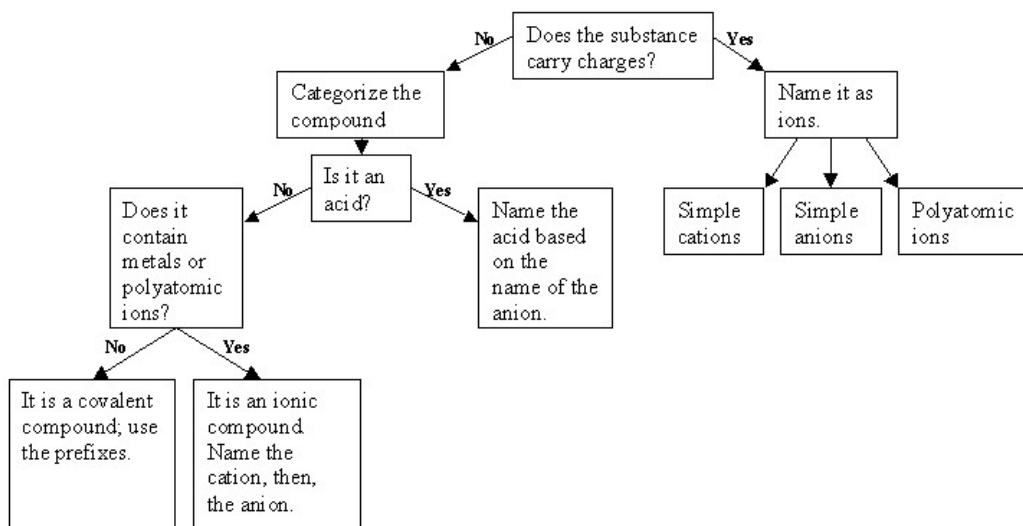
Strong bases end in OH, so they are all called _____ hydroxide.

Simply state the cation and then say hydroxide.



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Flowchart of naming compounds (copy this to your notes):



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Remember: Always determine what TYPE of compound you have first, only then will you be able to name it.

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